

<i>Positive Ions (Cations)</i>		<i>Negative Ions (Anions)</i>	
Name of Ion	Ion	Name of Ion	Ion
Aluminum	Al ⁺³	Bromide	Br ⁻¹
Ammonium	(NH ₄) ⁺¹	Carbonate	(CO ₃) ⁻²
Barium	Ba ⁺²	Hydrogen Carbonate <i>bicarbonate</i>	(HCO ₃) ⁻¹
Calcium	Ca ⁺²	Chlorate	(ClO ₃) ⁻¹
Cobalt	Co ⁺²	Chloride	Cl ⁻¹
Copper(I) <i>cuprous</i>	Cu ⁺¹	Chlorite	(ClO ₂) ⁻¹
Copper(II) <i>cupric</i>	Cu ⁺²	Chromate	(CrO ₄) ⁻²
Hydrogen	H ⁺¹	Fluoride	F ⁻¹
Iron(II) <i>ferrous</i>	Fe ⁺²	Hydroxide	(OH) ⁻¹
Iron(III) <i>ferric</i>	Fe ⁺³	Hypochlorite	(ClO) ⁻¹
Lead	Pb ⁺²	Iodide	I ⁻¹
Lithium	Li ⁺¹	Nitrate	(NO ₃) ⁻¹
Magnesium	Mg ⁺²	Nitrite	(NO ₂) ⁻¹
Mercury (I) <i>mercurous</i>	Hg ₂ ⁺²	Oxide	O ⁻²
Mercury (II) <i>mercuric</i>	Hg ⁺²	Perchlorate	(ClO ₄) ⁻¹
Nickel	Ni ⁺²	Permanganate	(MnO ₄) ⁻¹
Potassium	K ⁺¹	Phosphate	(PO ₄) ⁻³
Rubidium	Rb ⁺¹	Phosphide	P ⁻³
Silver	Ag ⁺¹	Sulfate	(SO ₄) ⁻²
Sodium	Na ⁺¹	Hydrogen Sulfate <i>bisulfate</i>	(HSO ₄) ⁻¹
Strontium	Sr ⁺²	Sulfide	S ⁻²
Tin (II) <i>stannous</i>	Sn ⁺²	Hydrogen Sulfide <i>bisulfide</i>	(HS) ⁻¹
Tin (IV) <i>stannic</i>	Sn ⁺⁴	Sulfite	(SO ₃) ⁻²
Zinc	Zn ⁺²	Hydrogen Sulfite <i>bisulfite</i>	(HSO ₃) ⁻¹

Name _____ Date _____ Per _____

Write the formulas for the following compounds

1.	Potassium Chloride	
2.	Magnesium Bromide	
3.	Aluminum Bromide	
4.	Aluminum Sulfide	
5.	Sodium Nitrate	
6.	Magnesium Hydroxide	
7.	Ammonium Chloride	
8.	Ferrous Sulfate	
9.	Ferric Chlorate	
10.	Ammonium Phosphate	
11.	Zinc Chromate	
12.	Iron (III) Oxide	
13.	Mercury (II) Phosphide	
14.	Tin (IV) Sulfate	
15.	Nickel Chromate	

Write the names for the following compounds

1.	PbO	
2.	K ₂ CO ₃	
3.	Ca ₃ (PO ₄) ₂	
4.	Al ₂ S ₃	
5.	Fe ₂ O ₃	
6.	NH ₄ NO ₃	
7.	NiCrO ₄	
8.	Fe(ClO ₂) ₂	
9.	NaHSO ₄	
10.	Sn(HS) ₂	